

## SNC2D – Chemistry Quiz – Answer Sheet

Name: \_\_\_\_\_

1	6	11	16	21	26	31	36
2	7	12	17	22	27	32	37
3	8	13	18	23	28	33	38
4	9	14	19	24	29	34	39
5	10	15	20	25	30	35	40

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<p>1. How many valence electrons does <b>fluorine</b> have?</p> <p>a) 1 b) 2 c) 7 d) 8 e) 9</p>	<p>6. Which of the following is the proper name for the chemical <b>NaBr</b>?</p> <p>a) sodium bromate b) sodium hydrobromate c) sodium bromine d) sodium bromide e) neighbour</p>
<p>2. Which of the following atoms would want to <b>lose 3 electrons</b> to form an ion?</p> <p>a) Nitrogen (N) b) Aluminum (Al) c) Oxygen (O) d) Calcium (Ca) e) Lithium (Li)</p>	<p>7. Which of the following is the proper name for the chemical <b>CaCO<sub>3</sub></b>?</p> <p>a) calcium monocarbon monoxide b) calcium carbonate c) calcium carbon oxide d) calcium carbonide e) calcium carbon trioxide</p>
<p>3. Which of the following describes how a negative ion is formed?</p> <p>a) atoms lose electrons b) atoms gain protons c) atoms lose protons d) atoms gain electrons e) atoms' neutrons become electrons</p>	<p>8. Which of the following is the proper name for the chemical <b>Al(CN)<sub>3</sub></b>?</p> <p>a) aluminum tricarbonitride b) aluminum carbon nitride c) aluminum thiocyanide d) aluminum cyanate e) aluminum cyanide</p>
<p>4. Selenium (Se) has 6 valence electrons. Which ion do you expect selenium to form?</p> <p>a) Se<sup>6+</sup> b) Se<sup>6-</sup> c) Se<sup>4+</sup> d) Se<sup>2+</sup> e) Se<sup>2-</sup></p>	<p>9. Which of the following is the proper name for the chemical <b>Fe<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub></b>?</p> <p>a) iron (III) sulfate b) iron (II) sulfate c) iron (III) sulfide d) iron (II) sulfide e) iron sulfide</p>
<p>5. Atoms of which of the following element prefer to become <b>cations</b>?</p> <p>a) calcium b) chlorine c) helium d) iodine e) oxygen</p>	<p>10. Which of the following is the proper name for the chemical <b>CuO</b>?</p> <p>a) copper oxide b) copper (I) oxide c) copper (II) oxide d) copper (I) monoxide e) copper (II) monoxide</p>

11. Which of the following is the proper name for the chemical **K<sub>3</sub>P**?

- a) potassium phosphide
- b) potassium phosphate
- c) krypton phosphorous
- d) potassium triphosphide
- e) tripotassium phosphide

12. Which of these chemical formulas correspond to **lithium hydride**?

- a) Li<sub>2</sub>H<sub>2</sub>
- b) LiH
- c) LiHe
- d) Li<sub>3</sub>H<sub>2</sub>
- e) LiH<sub>2</sub>

13. Which of these chemical formulas correspond to **lead (IV) sulfide**?

- a) Pb<sub>2</sub>S<sub>4</sub>
- b) Pb<sub>4</sub>S<sub>2</sub>
- c) Pb<sup>4+</sup>S<sup>2-</sup>
- d) PbS<sub>2</sub>
- e) Pb<sub>2</sub>S

14. Which of these chemical formulas correspond to **ammonium phosphate**?

- a) (NH<sub>4</sub>)<sub>3</sub>PO<sub>4</sub>
- b) NH<sub>4</sub>PO<sub>4</sub>
- c) NH<sub>4</sub>(PO<sub>4</sub>)<sub>3</sub>
- d) (NH<sub>4</sub>)(PO<sub>4</sub>)
- e) NP(OH)<sub>4</sub>

15. Which of these chemical formulas correspond to **cobalt (II) permanganate**?

- a) CoMnO<sub>4</sub>
- b) Co(MnO<sub>4</sub>)<sub>2</sub>
- c) Co(MnO<sub>2</sub>)<sub>2</sub>
- d) Co(Mn<sub>2</sub>O<sub>8</sub>)
- e) Co<sub>2</sub>MnO<sub>4</sub>

16. Which of the following is the proper name for the chemical **sodium nitride**?

- a) NaN<sub>3</sub>
- b) NaNO<sub>3</sub>
- c) NaNO
- d) Na(NO<sub>3</sub>)<sub>3</sub>
- e) Na<sub>3</sub>N

17. Which of the following does **not** characterize ionic compounds?

- a) Usually solid
- b) Conducts electricity as a solid
- c) Conducts electricity when dissolved in water
- d) High melting point
- e) Soluble in water

18. What is the proper name for **N<sub>2</sub>O<sub>5</sub>**?

- a) nitrogen (II) oxide (V)
- b) nitrogen (II) oxide
- c) dinitrogen pentaoxygen
- d) tetranitrogen decaoxygen
- e) dinitrogen pentaoxide

19. What is the proper name for **SCl<sub>2</sub>**?

- a) sulfur chloride
- b) monosulfur chloride
- c) monosulfur dichloride
- d) sulfur dichloride
- e) scandium iodide

20. Which of the following is **tetraphosphorous decaoxide**?

- a) P<sub>4</sub>O<sub>10</sub>
- b) P<sub>2</sub>O<sub>5</sub>
- c) P<sub>10</sub>O<sub>4</sub>
- d) P<sub>5</sub>O<sub>2</sub>
- e) PO<sub>2.5</sub>

21.  $Ca + AgNO_3 \rightarrow Ca(NO_3)_2 + Ag$  is an example of which type of reaction?

- a) synthesis
- b) decomposition
- c) single displacement
- d) double displacement
- e) combustion

22.  $2Na + Cl_2 \rightarrow 2NaCl$  is an example of which type of reaction?

- a) synthesis
- b) decomposition
- c) single displacement
- d) double displacement
- e) combustion

23.  $2NaCl + Pb(NO_3)_2 \rightarrow 2NaNO_3 + PbCl_2$  is an example of which type of reaction?

- a) synthesis
- b) decomposition
- c) single displacement
- d) double displacement
- e) combustion

24. Which of the following is exactly the same as  $O_2F_2$ ?

- a) OF
- b)  $O_3F_3$
- c)  $O_4F_4$
- d) all of the above
- e) none of the above

25.  $2H_2O_2 \rightarrow O_2 + 2H_2O$  is an example of which type of reaction?

- a) synthesis
- b) decomposition
- c) single displacement
- d) double displacement
- e) combustion

26. Which of the following will **not** replace aluminum in a single displacement reaction?

- a) zinc
- b) magnesium
- c) sodium
- d) lithium
- e) barium

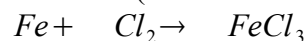
27. Which of the following elements will replace cobalt in a single displacement reaction?

- a) nickel
- b) magnesium
- c) hydrogen
- d) (a) and (b)
- e) (a), (b) and (c)

28.  $4C_2H_5 + 13O_2 \rightarrow 8CO_2 + 10H_2O$  is an example of which type of reaction?

- a) synthesis
- b) decomposition
- c) single displacement
- d) double displacement
- e) combustion

29. Balance this reaction (no fractions allowed):



What is the coefficient on  $Cl_2$ ?

- a) 1
- b) 2
- c) 3
- d) 4
- e) 5

30. How many atoms of hydrogen are in ammonium phosphate,  $(NH_4)_3PO_4$ ?

- a) 4
- b) 3
- c) 7
- d) 12
- e) 43

31. Which of the following is **hydrofluoric acid**?

- a)  $\text{H}_2\text{F}$
- b)  $\text{HFO}_3$
- c)  $\text{HF}$
- d)  $\text{H}_3\text{F}_2$
- e)  $\text{H}_2\text{OF}$

32. Which of the following is **nitric acid**?

- a)  $\text{H}_3\text{N}$
- b)  $\text{HONO}$
- c)  $\text{HN}$
- d)  $\text{HNO}_3$
- e)  $\text{NH}_3$

33. Which of the following is not a **base**?

- a)  $\text{NaOH}$
- b)  $\text{Ca}(\text{OH})_2$
- c)  $\text{Al}(\text{OH})_3$
- d)  $\text{LiOH}$
- e)  $\text{CH}_3\text{COOH}$

34. Which of the following best represents the pH of a base?

- a) higher than 14
- b) lower than 14
- c) higher than 7
- d) lower than 7
- e) higher than 0

35. What is the only material in the world to have a pH of exactly 7?

- a) tap water
- b) spring water
- c) distilled water
- d) pool water, after you add pH pellets
- e) urine

36. What are the products of this neutralization reaction:  $\text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow$

- a)  $\text{H}_2\text{O}$  and  $\text{NaSO}_4$
- b)  $\text{H}_2\text{OH}$  and  $\text{NaSO}_4$
- c)  $\text{NaH}_2$  and  $\text{OHSO}_4$  (also known as  $\text{HSO}_5$ )
- d)  $\text{H}_2\text{O}$ ,  $\text{Na}$  and  $\text{SO}_4$
- e)  $\text{H}_2\text{O}$  and  $\text{Na}_2\text{SO}_4$

37. If you have heartburn, which is where  $\text{HCl}$  in your stomach splashes up into your esophagus, which of the following should you eat/drink to soothe yourself?

- a)  $\text{H}_2\text{SO}_4$
- b)  $\text{CH}_3\text{COOH}$
- c)  $\text{Mg}(\text{OH})_2$
- d)  $\text{NaCl}$
- e) jalapeños!

38. It's not on your list, but which of the following is **chromium (VI) fluoride**?

- a)  $\text{Cr}_6\text{F}$
- b)  $\text{CrF}_6$
- c)  $\text{CrF}_3$
- d)  $\text{CrF}_2$
- e)  $\text{CrF}$

39. Which of the following does **not** dissolve in water?

- a) sodium fluoride
- b) calcium nitrate
- c) ammonium chloride
- d) barium carbonate
- e) silver chlorate

40. Which of the following **does** dissolve in water?

- a)  $\text{CuSO}_4$
- b)  $\text{AlF}_3$
- c)  $\text{Ca}_3(\text{PO}_4)_2$
- d)  $\text{AgI}$
- e)  $\text{ZnS}$